

# *The* **ULTIMATE** *New York City* **TRAVEL GUIDE**

BY: @NICOLE\_THENOMAD



# Visitors Guide To New York City

**L Cohen**



## **Visitors Guide To New York City:**

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web in the first semester lab course you may have determined the temperature dependence of the keto enol equilibrium constants of acetylacetone 2 4 pentanedione and ethyl acetoacetate ethyl 3 oxobutanoate in chloroform  $d_1$   $cdcl_3$  and methanol  $d_4$   $cd_3od$  using 1h nmr 1 4 see figure 1

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by integrating the enol and keto signal areas table 1 summarizes the equilibrium constant and percent enol concentration values calculated for the different combinations of enol to keto ratio note that

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web the relative amounts of products and reactants at equilibrium is characterized by the equilibrium constant for this reaction  $K_{eq}$  is defined as  $\frac{[enol]_{eq}}{[keto]_{eq}}$  where the subscript eq indicates that all concentrations are measured at equilibrium

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web because the system exists essentially completely as the thiol isomer a carbon acid acidity constant for ionization starting with the thio keto form as the initial state  $K_j$  could not be measured and a keto enol equilibrium constant  $K_{te}$  could not be determined

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web experimental methods detailed experimental procedure used for this experiment is given in the chem 417 laboratory manual the experiment 1 nmr determination of keto enol equilibrium constant each student prepared 4 tubes with sample solutions and two nmr tubes with solvent only one with deuterated chloroform and one with the methanol

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web using nmr spectroscopy the equilibrium values for different ketones which undergo tautomerization were found by determining relative peak values for each species the value for acetylacetone was found to be  $K_c = 6.622$  and for ethyl acetoacetate was  $K_c = 7.510$

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web nmr determination of keto enol equilibrium constants the nmr spectrum below is for acetylacetone with  $CDCl_3$  as solvent the photos depict the complete nmr spectrum as well more zoomed in areas of each peak demonstrate how to calculate the equilibrium constant for keto to enol conversion  $K_c$