

CLASS: SS 2 SUBJECT: CHEMISTRY TIME: 2 HOURS

SECTION A – OBJECTIVES

Instructions: Answer all questions

1. Reaction represented by the equation $\text{MnO}_4^- + 5\text{e}^- \longrightarrow \text{Mn}^{2+}$ is known as A. Hydrolysis B. Oxidation C. Reduction D. Fermentation
2. The reaction represented by the equation $\text{NaOH}_{(\text{aq})} + \text{HCl}_{(\text{aq})} \longrightarrow \text{NaCl} + \text{H}_2\text{O}$ is an A. double decomposition reaction B. Is a neutralization reaction C. Is reversible D. Is usually catalyzed
3. Which of the following does NOT affect the rate of a chemical reaction? A. Concentration of the reactants B. Addition or presence of a catalyst C. size of the reacting particles D. The enthalpy changes for the reaction
4. Given that 32.0g of sulphur contains 6.02×10^{23} sulphur atoms, how many atoms are there in 2.70g of aluminum? A. 6.02×10^{23} B. 3.01×10^{23} C. 6.02×10^{22} D. 5.08×10^{22}
5. A salt which loses mass when exposed to air is A. Hygroscopic B. Deliquescent C. Efflorescent D. Fluorescent
6. The following are heavy chemicals EXCEPT A. tetraoxosulphate(vi) acid B. Soda C. Sodium trioxocarbonate(iv) D. Dyes
7. Which of the following correctly explains why iodine crystals change directly into the gaseous state when heated? A. Iodine crystals sublime B. Iodine atoms are held together by strong forces C. Iodine molecules are held together by strong forces D. Iodine molecules are held together by weak forces
8. The major raw material in the plastic industry is A. Ethanol B. Sulphur C. Methylethanoate D. Ethene
9. When an element can exist in two or more forms in the same physical state, the element is said to exhibit A. Isotopy B. Isomerism C. Allotropy D. Hygroscopy
10. What is the concentration, in mol dm^{-3} , of a solution containing 0.10g of sodium hydroxide per 50cm^3 of solution? A. 0.05 B. 0.1 C. 0.5 D. 2
11. The following are s-block elements EXCEPT A. Lithium B. Magnesium C. Aluminum D. Barium
12. The electronic configurations of two atoms X and Y are as follows;
13. X: $1\text{s}^2 2\text{s}^2 2\text{p}^6 3\text{s}^1$, Y: $1\text{s}^2 2\text{s}^2 2\text{p}^6 3\text{s}^2 3\text{p}^6 4\text{s}^2$. Which of the statements below is true of the position of X and Y in the periodic table?
 - A. X belongs to group 1, Y belongs to period 2
 - B. X belongs to group 1, Y belongs to period 4
 - C. X belongs to group 2, Y belongs to period 1
 - D. X belongs to group 3, Y belongs to period 2
14. The oxidation number of chloride is +1 in A. KClO_3 B. Cl_2O_7 C. ZnCl_2 D. NaClO
15. Which of the following separation is routinely applied in the petroleum industry A. Filtration B. Chromatography C. Evaporation D. Fractional distillation
16. What does X stand for in the following expression? $X = -\log_{10} [\text{H}^+]$
 - A. Enthalpy change, H
 - B. Equilibrium constant, K
 - C. Solubility product
 - D. Degree of acidity, pH
17. Which of the following indicates correctly the increasing order of oxidizing power of the halogens? A. $\text{F} < \text{Cl} < \text{Br} < \text{I}$ B. $\text{I} < \text{Br} < \text{Cl} < \text{F}$ C. $\text{I} < \text{Cl} < \text{Br} < \text{F}$ D. $\text{Br} < \text{I} < \text{Cl} < \text{F}$
18. Which of the following compounds has the highest ionic character? A. $\text{KCl}_{(\text{s})}$ B. $\text{BeCl}_{2(\text{g})}$ C. $\text{HCl}_{(\text{g})}$ D. $\text{AlCl}_{3(\text{g})}$

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