

Exercises

14.1-14.3 Temperature, Internal energy and states of matter

- Molecules (of the liquid) are in random / haphazard motion (AM)
 - Molecules (of liquid) are smaller than pollen grains.
 - increase the temperature (of the liquid)
- The sum of (the random distribution of) the KE and PE of (the) molecules
- No change in KE

Because temperature is constant (during melting)

PE of (the molecules) increases (during melting)

The internal energy increases.

14.4 Specific Heat Capacity

- 35 700 J
- 21 000 J/g°C
- 0.375 kg
- 66 J/s
- Electrical energy supplied = Heat taken in by water

$$I \times V \times t = m \times c \Delta \theta$$

$$2300 \times I \times 2 = 2.5 \times 4200 \times 80$$

$$I = 380 \text{ A [1]}$$
 - The kinetic energy of the particle is increasing. [1]
- $$E = 3 \times 60 = 180 \text{ J [1]}$$

Thermal Energy And Thermodynamics Test Review Answers

D Siedentop



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