

Practice Problems

Friday, 27 August 2021 3:15 PM

Heat Capacity and Calorimetry

[1]

How many calories are required to heat each of the following from 15°C to 65°C: (a) 1.0 g water; (b) 5.0 g Pyrex[®] glass; (c) 20 g platinum? (The specific heat of the glass is 0.20 cal/g · K; of platinum, 0.032 cal/g · K.)

Ans. (a) 50 cal; (b) 50 cal; (c) 32 cal

[2]

The ability to convert between calories and joules is extremely important. Express the answers in Problem 7.19 in joules and in kJ.

Ans. (a and b) 209.2 J and 0.2092 kJ; (c) 133.9 J and 0.1339 kJ

[3]

The combustion of 5.00 g of coke raised the temperature of water of 1.00 kg of water from 10°C to 47°C. Calculate the heating value of coke.

Ans. 7.4 kcal/g or 31 kJ/g

[4]

Assuming that 50% of the heat is useful, how many kilograms of water at 15°C can be heated to 95°C by burning 200 L of methane, CH₄, measured at S.T.P.? The heat of combustion for methane is 891 kJ/mol.

Ans. 12 kg water

[5]

The heat of combustion of ethane gas, C₂H₆, is 1561 kJ/mol. Assuming that 60% of the heat is useful, how many liters of ethane (S.T.P.) must be burned to supply enough heat to convert 50 kg of water at 10°C to steam at 100°C?

Ans. 3150 L

[6]

How much heat is required to change 10 g of ice at 0°C to steam at 100°C? The heat of vaporization of water at 100°C is 540 cal/g (2259 J/g).

Ans. 7.2 kcal or 30 kJ

[7]

A boiler that supplies a steam turbine used in an electric power plant must raise the temperature of water from the ambient temperature passing through a physical state change and then to live (very hot) steam. How much energy (kcal and kJ) would be required to raise 15,000 L of water from 25°C to steam at 175°C? (steam requires 0.5 cal/g · K)?

Ans. 9,800 kcal or 41,000 kJ

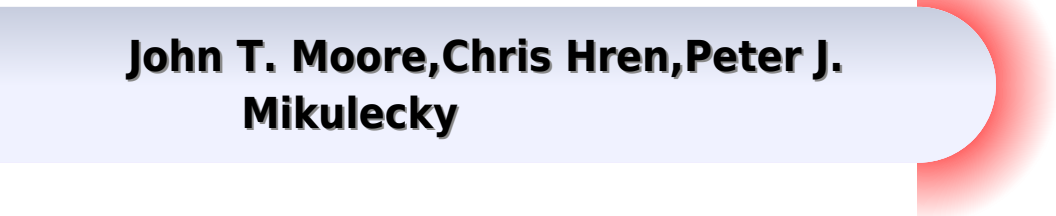
[8]

A Styrofoam[™] coffee cup serves as an inexpensive calorimeter for measurements that do not require high accuracy. One gram of KCl(s) was added to 25.0 ml of water in such a cup at 24.33°C. It dissolved completely and quickly with gentle stirring. The minimum temperature reached was 22.12°C. Estimate the ΔH° of the heat of solution of KCl in kJ/mol. You may assume the solution has the same heat capacity as water and that the heat capacity of the cup and thermometer need not be considered.

Ans. $\Delta H^\circ_{\text{solution}} = +17.2 \text{ kJ/mol}$

Thermochemistry Practice Calculations Answers

**John T. Moore, Chris Hren, Peter J.
Mikulecky**



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Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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Introductory Chemistry Problems: The Chemistry And Physics Behind The Numbers Michael E Green, 2025-07-18 The essential point of the book is to show students in Introductory Chemistry classes how to obtain answers that are at least

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