

Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone, CaCO_3 , to lime, CaO , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{ CaCO}_3(s) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CaO}(s) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CO}_2(g) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4. $\text{CH}_4(g) + 3 \text{Cl}_2(g) \rightarrow \text{CHCl}_3(l) + 3 \text{HCl}(g) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide, CH_2N_2 , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: $\Delta H^\circ_f \text{ NO}(g) = 90.3 \text{ kJ/mol}$ and $\Delta H^\circ_f \text{ H}_2\text{O}(g) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction: $2 \text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + 2 \text{Fe}(s) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe_2O_3 , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

Thermochemistry Practice Problem Answers

Donald A. McQuarrie, Stanley Gill



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Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *Ebbing* Darrell D. Ebbing,1993 **Cracking the GRE Chemistry Subject Test** Princeton Review (Firm),2005 The GRE subject tests are among the most difficult standardized exams Rather than testing general problem solving skills they require highly specialized knowledge The experts at The Princeton Review have thoroughly research each subject test to provide students with the most thorough up to date information available Students don t need to relearn the entire histories of their fields just what they need to know to earn high scores on the exams Each guide includes one full length practice exam complete with comprehensive explanations for every solution

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