#### Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process

 Calculate the heat evolved in the combustion of ethyl alcohol, ΔH<sub>comb</sub> C<sub>2</sub>H<sub>5</sub> OH<sub>18</sub> + 3 O<sub>2 los</sub> → 2 CO<sub>2 los</sub> + 3 H<sub>2</sub>O<sub>18</sub> ΔH = ?

2 Consense + 3 H <sub>2 (q)</sub> + ½ O <sub>2 (q)</sub>	→ C <sub>2</sub> H <sub>5</sub> OH <sub>8</sub>	$\Delta H = -277.7  kJ$
Cgraphes + O2 kg → CO2 kg		$\Delta H = -393.5 \text{ kJ}$
H <sub>2 lo0</sub> + ½ O <sub>2 lot</sub> → H <sub>2</sub> O <sub>II</sub>		ΔH = -285.8 kJ
		$\Delta H = -1366.7 \text{ kJ}$

Calculate the heat required to decompose limestone, CaCO<sub>3</sub>, to lime, CaO, and carbon dioxide, with all substances at standard conditions.
CaCO<sub>3 N</sub> → CaO<sub>N</sub> + CO<sub>2 N</sub>

ΔH\*<sub>1</sub> CaCO<sub>3 |q|</sub> = -1206.9 kJ/mol ΔH\*<sub>1</sub> CaO<sub>3 |q|</sub> = -635.1 kJ/mol ΔH\*<sub>1</sub> CO<sub>2 |q|</sub> = -393.5 kJ/mol

 $\Delta H = 178.3 \text{ kJ}$ 

-46.1 kJ/mel

4.  $CH_{E(g)} + 3Cl_{E(g)} \rightarrow CHCl_{D(g)} + 3HCl_{(g)} \Delta H = ?$ 

 Cyanamide, CH<sub>2</sub>N<sub>2</sub>, is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:

The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.
 4 NH<sub>3 kgi</sub> + 5 O<sub>2 kgi</sub> → 4 NO kgi + 6 H<sub>2</sub>O<sub>3gi</sub>
 Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: ΔH\*<sub>1</sub> NO kgi = 90.3 kJ/mol and ΔH\*<sub>1</sub> H<sub>2</sub>O kgi = -241.8 kJ/mol

7. The thermite reaction:  $2 \text{ Al}_{|\mathbb{N}|} + \text{Fe}_2 O_{3 |\mathbb{N}|} \rightarrow \text{Al}_2 O_{3 |\mathbb{N}|} + 2 \text{ Fe}_{|\mathbb{N}|} + 352 \text{ kJ}$ 

If you begin with 10.0 g of Al and excess Fe<sub>2</sub>O<sub>3</sub>, how many kJ of heat are evolved at constant pressure?

# **Thermochemistry Practice Calculation Unit 12**

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