

First law of thermodynamics for various processes:

- i. **Isothermal process:** In such process, temperature is constant, internal energy is constant, hence $\Delta U = 0$

$$\text{Therefore } \Delta U = Q + W$$

$$0 = Q + W$$

$$\text{Or } W = -Q$$

- ii. **Adiabatic process:** In this process, there is no exchange of heat between system and its surroundings

$$\text{i.e. } Q = 0, \text{ then}$$

$$-\Delta U = -W$$

- iii. **Isochoric process:**

we know,

$$W = -P_{\text{ext}} \Delta V$$

$$\Delta U = Q + W$$

$$\Delta U = Q - P_{\text{ext}} \Delta V$$

In isochoric process

$$\Delta V = 0$$

$$\text{Therefore } \Delta U = Q_v$$

U is state function, Q_v is also a state function.

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Thermodynamics 12th Std

Lakhmir Singh & Manjit Kaur



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