

Name: Key

Date: _____

Block: _____

Molarity and Dilutions Practice - HN**Part 1: Molarity**

$$M = \frac{\text{moles of solute}}{\text{Volume of Solution (L)}} = \frac{\text{Mols}}{\text{L}}$$

1. What is the molarity of a 0.30 liter solution containing 0.50 moles of sodium chloride, NaCl?

$$M = \frac{.5 \text{ mol}}{.3 \text{ L}}$$

$$M = \boxed{1.67 \text{ M}}$$

2. Calculate the molarity of 0.289 moles of Iron (III) Chloride, FeCl_3 , dissolved in 120 mL of solution.

$$\textcircled{1} \quad 120 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .12 \text{ L}$$

mL $\rightarrow$ L

$$\textcircled{2} \quad M = \frac{.289 \text{ mol}}{.12 \text{ L}} = \boxed{2.41 \text{ M}}$$

3. What is the molarity of 0.5 grams of sodium chloride, NaCl, dissolved to make 50 mL of solution?

$$\textcircled{1} \quad .5 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}} = .009 \text{ mol}$$

g $\rightarrow$ mol

$$\textcircled{2} \quad 50 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .05 \text{ L}$$

$$\textcircled{3} \quad M = \frac{.009 \text{ mol}}{.05 \text{ L}} = \boxed{.18 \text{ M}}$$

4. Calculate the molarity of 734 grams of lithium sulfate, Li_2SO_4 , dissolved in 2,500 mL of solution.

$$\textcircled{1} \quad 734 \text{ g} \times \frac{1 \text{ mol}}{109.95 \text{ g}} = 6.68 \text{ mol}$$

$$\textcircled{2} \quad 2500 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 2.5 \text{ L}$$

$$\textcircled{3} \quad M = \frac{6.68 \text{ mol}}{2.5 \text{ L}} = \boxed{2.67 \text{ M}}$$

5. How many grams of Ca(OH)_2 are needed to produce 500 mL of a 1.66M Ca(OH)_2 solution?

$$\textcircled{1} \quad 500 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .5 \text{ L}$$

$$\textcircled{2} \quad 1.66 \text{ M} = \frac{\text{mol}}{.5 \text{ L}}$$

mol = .83 mol

$$\textcircled{3} \quad .83 \text{ mol} \times \frac{74.1 \text{ g}}{1 \text{ mol}} = \boxed{61.5 \text{ g Ca(OH)}_2}$$

6. How many grams of KOH are required to make 125 mL of a 0.50M solution?

$$\textcircled{1} \quad 125 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .125 \text{ L}$$

$$\textcircled{2} \quad .5 \text{ M} = \frac{\text{mol}}{.125}$$

mol = .063 mol

$$\textcircled{3} \quad .063 \text{ mol} \times \frac{56.105 \text{ g}}{1 \text{ mol}} = \boxed{3.53 \text{ g}}$$

7. How many grams of Na_2SO_4 are required to make 575 mL of a 0.60M solution?

$$\textcircled{1} \quad 575 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .575 \text{ L}$$

$$\textcircled{2} \quad .6 \text{ M} = \frac{\text{mol}}{.575 \text{ L}}$$

mol = .345

$$\textcircled{3} \quad .345 \text{ mol} \times \frac{142.04 \text{ g}}{1 \text{ mol}} = \boxed{49.0 \text{ g}}$$

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Table of Contents Solutions Molarity And Dilution Practice Answer Key

1. Understanding the eBook Solutions Molarity And Dilution Practice Answer Key
 - The Rise of Digital Reading Solutions Molarity And Dilution Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Molarity And Dilution Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Molarity And Dilution Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Molarity And Dilution Practice Answer Key
 - Personalized Recommendations
 - Solutions Molarity And Dilution Practice Answer Key User Reviews and Ratings
 - Solutions Molarity And Dilution Practice Answer Key and Bestseller Lists

5. Accessing Solutions Molarity And Dilution Practice Answer Key Free and Paid eBooks
 - Solutions Molarity And Dilution Practice Answer Key Public Domain eBooks
 - Solutions Molarity And Dilution Practice Answer Key eBook Subscription Services
 - Solutions Molarity And Dilution Practice Answer Key Budget-Friendly Options
6. Navigating Solutions Molarity And Dilution Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Solutions Molarity And Dilution Practice Answer Key Compatibility with Devices
 - Solutions Molarity And Dilution Practice Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solutions Molarity And Dilution Practice Answer Key
 - Highlighting and Note-Taking Solutions Molarity And Dilution Practice Answer Key
 - Interactive Elements Solutions Molarity And Dilution Practice Answer Key
8. Staying Engaged with Solutions Molarity And Dilution Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Molarity And Dilution Practice Answer Key
9. Balancing eBooks and Physical Books Solutions Molarity And Dilution Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Molarity And Dilution Practice Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solutions Molarity And Dilution Practice Answer Key
 - Setting Reading Goals Solutions Molarity And Dilution Practice Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solutions Molarity And Dilution Practice Answer Key
 - Fact-Checking eBook Content of Solutions Molarity And Dilution Practice Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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