Subject:- Chemistry

Chemical Thermodynamics and Energetics

Practice Problems-II

Q1) Three moles of an ideal gas are expanded isothermally from a volume of 300 cm³ to 2.5 L at 300 K against a pressure of 1.9 atm. Calculate the work done in L atm and joules.

Soln:-
$$V_1 = 300$$
 c.c. $= 0.3$ L, $V_2 = 2.5$ L, $T = 300$ K, $P_{ex} = 1.9$ atm $W = -P_{ex}\Delta V = -P_{ex}(V_2 - V_1)$ $= -1.9(2.5 - 0.3) = -1.9 \times 2.2$ $\therefore W = -4.18$ L atm and $W = -4.18 \times 101.3 = -423.4$ J

Q2) One mole of an ideal gas is compressed from $500~\rm cm^3$ against a constant pressure of $1.216\times 10^5~\rm Pa$. The work involved in the process is $36.50~\rm J$. Calculate the final volume.

Soln:-
$$V_1 = 500 \text{ c.c.} = 0.5 \text{ dm}^3$$
, $P_{ex} = 1.216 \times 10^5 \text{ Pa} = 1.216 \times 10^2 \text{ kPa}$, $W = 36.50 \text{ J}$ $W = -P_{ex}\Delta V$

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