

Chapter 16 Redox chemistry and corrosion

Q1.

Identify each of the following half equations as involving either oxidation or reduction:

- a $\text{Na(s)} \rightarrow \text{Na}^+(\text{aq}) + \text{e}^-$
- b $\text{Cl}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{Cl}^-(\text{aq})$
- c $\text{S(s)} + 2\text{e}^- \rightarrow \text{S}^{2-}(\text{aq})$
- d $\text{Zn(s)} \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$

A1.

- a oxidation
- b reduction
- c reduction
- d oxidation

Q2.

Balance the following half equations and then identify each as representing either an oxidation or a reduction reaction:

- a $\text{Fe(s)} \rightarrow \text{Fe}^{3+}(\text{aq})$
- b $\text{K(s)} \rightarrow \text{K}^+(\text{aq})$
- c $\text{F}_2(\text{g}) \rightarrow \text{F}^-(\text{aq})$
- d $\text{O}_2(\text{g}) \rightarrow \text{O}^{2-}(\text{aq})$

A2.

- a $\text{Fe(s)} \rightarrow \text{Fe}^{3+}(\text{aq}) + 3\text{e}^-$ oxidation
- b $\text{K(s)} \rightarrow \text{K}^+(\text{aq}) + \text{e}^-$ oxidation
- c $\text{F}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{F}^-(\text{aq})$ reduction
- d $\text{O}_2(\text{g}) + 4\text{e}^- \rightarrow 2\text{O}^{2-}(\text{aq})$ reduction

Q3.

Iron reacts with hydrochloric acid according to the ionic equation:



- a What has been oxidised in this reaction? What is the product?
- b Write a half equation for the oxidation reaction.
- c Identify the oxidant.
- d What has been reduced in this reaction? What is the product?
- e Write a half equation for the reduction reaction.
- f Identify the reductant.

A3.

- a Fe(s) has been oxidised to $\text{Fe}^{2+}(\text{aq})$
- b $\text{Fe(s)} \rightarrow \text{Fe}^{2+}(\text{aq}) + 2\text{e}^-$
- c $\text{H}^+(\text{aq})$
- d $\text{H}^+(\text{aq})$ has been reduced to H_2
- e $2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$
- f Fe(s)

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