

Hydrogen $2H$

Oxygen O_2

Oxygen is 16x larger than hydrogen

You shouldn't conclude that chemists were able to determine the molar masses of all the elements using this technique. Measurements of the density of the gaseous phase of many of the elements are difficult, if not impossible. However, we are going to see that chemists could use another tool - the *percent composition of compounds* to determine molar masses.

Relative Mass From Compounds

Many substances combine with oxygen to form a type of compound called an oxide. In the previous unit we saw that such combinations often occur in multiple proportions. John Dalton made the assumption that the lowest ratio was a 1:1 combination of elements. For now, we will make a similar assumption. We may have to re-examine this assumption later.

In Unit 4 you examined % composition data of several compounds. For example, 100 g of the simplest oxide of carbon contains 42.9 g of C and 57.1 g of O. To be able to compare masses of elements to the mass of oxygen, it is important to use the same mass of oxygen in all of the ratios. Using a proportion, one can show that 75.0 g of C combines with 100 g of oxygen.

$$\textcircled{A} \quad \frac{42.9 \text{ g C}}{57.1 \text{ g O}} = \frac{75.0 \text{ g C}}{100 \text{ g O}}$$

$$\textcircled{B} \quad \frac{27.3 \text{ g C}}{72.7 \text{ g O}} = \frac{75.0 \text{ g C}}{100 \text{ g O}}$$

$$\frac{57.1 \text{ g O}}{42.9 \text{ g C}} = \frac{1.33 \text{ O}_2}{1 \text{ H}}$$

$$\frac{72.7 \text{ g O}}{27.3 \text{ g C}} = 2.66 \text{ O}_2$$

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CL Gary



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Table of Contents U5 Wsv3 0 Answers Modeling

1. Understanding the eBook U5 Wsv3 0 Answers Modeling
 - The Rise of Digital Reading U5 Wsv3 0 Answers Modeling
 - Advantages of eBooks Over Traditional Books
2. Identifying U5 Wsv3 0 Answers Modeling
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an U5 Wsv3 0 Answers Modeling
 - User-Friendly Interface
4. Exploring eBook Recommendations from U5 Wsv3 0 Answers Modeling
 - Personalized Recommendations
 - U5 Wsv3 0 Answers Modeling User Reviews and Ratings
 - U5 Wsv3 0 Answers Modeling and Bestseller Lists
5. Accessing U5 Wsv3 0 Answers Modeling Free and Paid eBooks

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 - U5 Wsv3 0 Answers Modeling eBook Subscription Services
 - U5 Wsv3 0 Answers Modeling Budget-Friendly Options
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 - U5 Wsv3 0 Answers Modeling Compatibility with Devices
 - U5 Wsv3 0 Answers Modeling Enhanced eBook Features
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 - Highlighting and Note-Taking U5 Wsv3 0 Answers Modeling
 - Interactive Elements U5 Wsv3 0 Answers Modeling
 8. Staying Engaged with U5 Wsv3 0 Answers Modeling
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers U5 Wsv3 0 Answers Modeling
 9. Balancing eBooks and Physical Books U5 Wsv3 0 Answers Modeling
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection U5 Wsv3 0 Answers Modeling
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine U5 Wsv3 0 Answers Modeling
 - Setting Reading Goals U5 Wsv3 0 Answers Modeling
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of U5 Wsv3 0 Answers Modeling
 - Fact-Checking eBook Content of U5 Wsv3 0 Answers Modeling
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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