

Now _____ Day _____ One _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your notebook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
---------------	----------	------------	------------	--------

When sharing of electron occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) p bond**.

In your notebook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form the elements such as hydrogen, nitrogen, and oxygen normally exist?
 - as single atoms
 - as molecules containing three atoms
 - as molecules containing two atoms
 - as molecules containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - three
- Bond length is the distance between
 - two molecules of the same substance
 - the nucleus of two attached atoms
 - the electron in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - the shorter the bond, the stronger the bond
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the lower the electron in it
 - the shorter the bond, the lower the bond dissociation energy

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49

Now _____ Day _____ One _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your notebook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

- Binary molecular compounds are generally composed of a metal and a nonmetal. **false**
- The second element in the formula of a binary compound is named using the suffix -ite. **false**
- The prefix zero indicates three atoms. **false**
- The prefix four indicates six atoms. **true**
- In naming the first element in a formula, the prefix mono- is not used. **true**
- For binary acids, the hydrogen part of the compound is named using the prefix hydro-. **true**
- An oxyacid contains only two elements. **false**
- If the name of the anion of an oxyacid ends in -ite, the acid name contains the suffix -ous. **false**

In your notebook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
1. $\text{C}_2\text{H}_4\text{O}_2$	a. hydrobromic acid
2. H_2SO_4	b. nitrogen trioxide
3. H_2CO_3	c. carbon monoxide
4. H_2O	d. nitrous acid
5. H_2SO_3	e. ammonia
6. H_2O_2	f. acetic acid
7. H_2SO_2	g. carbonic acid
8. HNO_3	h. bromic acid
9. HNO_2	i. carbon dioxide

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50

Study Guide For Content Mastery Answer Key Covalent Bonding

Martin Silberberg, Libby Weberg



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1. Understanding the eBook Study Guide For Content Mastery Answer Key Covalent Bonding
 - The Rise of Digital Reading Study Guide For Content Mastery Answer Key Covalent Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Answer Key Covalent Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Answer Key Covalent Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Answer Key Covalent Bonding
 - Personalized Recommendations
 - Study Guide For Content Mastery Answer Key Covalent Bonding User Reviews and Ratings
 - Study Guide For Content Mastery Answer Key Covalent Bonding and Bestseller Lists
5. Accessing Study Guide For Content Mastery Answer Key Covalent Bonding Free and Paid eBooks
 - Study Guide For Content Mastery Answer Key Covalent Bonding Public Domain eBooks
 - Study Guide For Content Mastery Answer Key Covalent Bonding eBook Subscription Services
 - Study Guide For Content Mastery Answer Key Covalent Bonding Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Study Guide For Content Mastery Answer Key Covalent Bonding Compatibility with Devices
 - Study Guide For Content Mastery Answer Key Covalent Bonding Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Study Guide For Content Mastery Answer Key Covalent Bonding
 - Highlighting and Note-Taking Study Guide For Content Mastery Answer Key Covalent Bonding
 - Interactive Elements Study Guide For Content Mastery Answer Key Covalent Bonding
8. Staying Engaged with Study Guide For Content Mastery Answer Key Covalent Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Answer Key Covalent Bonding
9. Balancing eBooks and Physical Books Study Guide For Content Mastery Answer Key Covalent Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Mastery Answer Key Covalent Bonding
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content Mastery Answer Key Covalent Bonding
 - Setting Reading Goals Study Guide For Content Mastery Answer Key Covalent Bonding
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide For Content Mastery Answer Key Covalent Bonding
 - Fact-Checking eBook Content of Study Guide For Content Mastery Answer Key Covalent Bonding
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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