

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of halogen bond: covalent and ionic.

• **Covalent bonds** are formed by the sharing of electrons. They are stronger than ionic bonds (polar) and weaker (dipolar). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of σ and π Bonds



Bond and Reaction Rates



Bond Energy



There are two types of halogen bond

Definition of halogen bond



Sigma (σ) and pi (π) bonds

Covalent bonds are classified as sigma (σ) and pi (π).

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **Pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two multiple bonds** consist of one sigma and π bonds.



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

values, covalent bonds can be polar and nonpolar.



• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar). For two atoms, forming a nonpolar covalent bond. Two important examples are the C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, we have a polar covalent bond. Most common examples are atoms for carbon, nitrogen, O, F and other halogens.

• A difference of 1.7 or higher can be high that the electrons are not shared (ionic) and we have ionic bond.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



Electronegative atoms are color-coded (blue) and electropositive atoms are color-coded (red).

• The dipole moment is always a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• δ^- (delta minus) indicates an electron-deficient atom.

• δ^+ (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar (even for dipole moments) are cancelled. A good example is carbon dioxide.



Cancel the vector or the other pointing point in the same way.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds. Double bonds are stronger than single bonds.

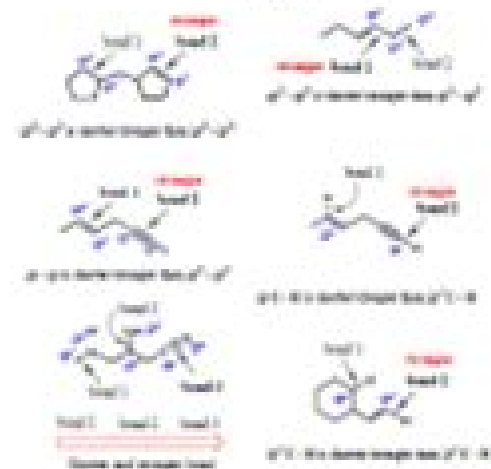


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	473 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Hybridization	25% s	33% s	50% s
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	473 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

Examples



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**Seyhan N. Ege,Roberta
Kleinman,Marjorie L. C. Carter**



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