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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_c is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant (K_p) For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_p is given by: $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
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Le Chatelier's Principle: The effect of changing conditions on equilibrium		
Concentration - Increasing the concentration of a reactant shifts the equilibrium to the right. - Increasing the concentration of a product shifts the equilibrium to the left.	Pressure - Increasing the pressure shifts the equilibrium to the side with fewer moles of gas. - Decreasing the pressure shifts the equilibrium to the side with more moles of gas.	Temperature - Increasing the temperature shifts the equilibrium to the endothermic direction. - Decreasing the temperature shifts the equilibrium to the exothermic direction.

Study Guide For Content Mastery Chemical Equilibrium

Julia Schneider



Study Guide For Content Mastery Chemical Equilibrium:

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homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04

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