

Gas Laws Study Guide

Test Format:

- Multiple choice and problems

Items that will be provided for the test (read important notes at the bottom):

- Gas law equations
- The value of R
- Pressure conversion factors

* I will provide the equations on the test. You are responsible for knowing which laws the equations describe (ex: I will give you $P_1 \times V_1 = P_2 \times V_2$ and you must know that this is Boyle's Law).

- What is the name of the model we currently use to explain the behavior of gases? **kinetic molecular theory (KMT)**
- Define the kinetic molecular theory (KMT). **Gases are in constant motion with elastic collisions**
- The KMT makes many assumptions about the Size, motion, and energy of gas particles.
- According to the KMT, do gas particles experience attractive or repulsive forces? Explain why or why not. **Neither- they are too far apart**
- Completely describe the motion of gas particles according to the KMT. **Constant, Random, Straight paths, different velocity**
- Define elastic collision. **No kinetic energy is lost- just transferred between particles**
- What determines the kinetic energy of a particle? **Mass and velocity**
- In a sample of a single gas, do all particles have the same kinetic energy? Explain why or why not. **No — they have a different velocity**
- Define temperature. **Measurement of kinetic energy**
- Explain why gases have a low density. **Lots of space between particles (large volume)**
- Why are gases compressible? **Lots of empty space between particles**
- Why do gases expand to fill a container? **No attractive or repulsive forces**
- Why can gas particles flow easily past each other? **No attractive forces**
- Define diffusion. **Movement of one material through another**
- Define effusion. **Gas escapes through tiny openings**
- Describe how diffusion depends on the mass of the particles. **Diffusion is dependent on mass, so less mass more diffusion**
- Use the KMT to explain why a balloon expands as you blow air into it. **Particles are in constant motion and collide with their surroundings expanding the balloon**

Boyle's Law: Answer the following questions about Boyle's Law.

- Describe Boyle's Law in your own words. **As pressure increases, volume decreases**
- What is the formula? **$P_1V_1 = P_2V_2$**
- What are the variables? What are the possible units used in the formula? **Pressure and volume, atm, kPa, mL, L**
- What is constant? **temperature**



- Give an example to help explain the law (you should have some type of visual of your example).
- What would a graph of pressure vs. volume of a gas look like according to Boyle's Law?

Charles' Law: Answer the following questions about Charles' Law.

- Describe Charles' Law in your own words. **Increase temperature, increase volume**
- What is the formula? **$V_1/T_1 = V_2/T_2$**
- What are the variables? What are the possible units used in the formula? **Temperature and volume (kelvin, mL, L)**
- What is constant? **Pressure**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Charles's Law (you should have some type of visual of your example).



- What would a graph of temperature vs. volume of a gas look like according to Charles's Law?

Gay-Lussac's Law: Answer the following questions about Gay-Lussac's Law.

- Describe Gay-Lussac's Law in your own words. **As pressure increases, temperature increases**
- What is the formula? **$P_1/T_1 = P_2/T_2$**
- What are the variables? What are the possible units used in the formula? **Pressure and temperature (kelvin, atm, kPa, torr)**
- What is constant? **volume**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Gay-Lussac's Law (you should have some type of visual of your example).

Study Guide Gas Laws

Karin Nielsen-Saines



Study Guide Gas Laws:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Physics for Scientists and Engineers Study Guide Todd Ruskell, 2007-06-08 The Study Guide provides students with key physical quantities and equations misconceptions to avoid questions and practice problems to gain further understanding of physics concepts and quizzes to test student knowledge of chapters

ASAP Chemistry: A Quick-Review Study Guide for the AP Exam The Princeton Review, 2019-02-12 Looking for sample exams practice questions and test taking strategies Check out our extended in depth AP chem prep guide

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General, Organic, and Biochemistry Study Guide Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian s Organic and Biochemistry

The Primary FRCA Structured Oral Examination Study Guide 1 Lara Wijayasiri, Kate McCombe, Amish Patel, 2010 This fully up to date book is designed specifically for candidates preparing for the Primary FRCA structured oral

examination incorporating the new exam structure and syllabus Sample questions accurately reflect the examination while model answers are systematically structured with definitions and classifications and illustrated with essential diagrams and graphs The books provide clear and concise explanations to key scientific concepts and problem based answers to clinical scenarios This first part contains questions on physiology and physics Packed with new guidelines fundamental topics that are poorly covered in other main texts and current hot topics this book and its companion The Primary FRCA Structured Oral Examination Study Guide 2 are the definitive revision aid to the Primary FRCA but will also be of value to candidates preparing for the basic science component of the Final FRCA as reference source for qualified anaesthetists and as a text for tutors preparing candidates for the structured oral examination SOE

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 Barron s Science 360 Chemistry is your complete go to guide for everything chemistry This

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Chemical Principles Student's Study Guide & Solutions Manual John Krenos,Joseph Potenza,Loretta Jones,Lynn Koplitz,Thomas Spence,2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

Chemical Principles Study Guide/Solutions Manual John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn,Rebecca Brewer,Rachel C. Lum,2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

Research in Education ,1974

REHS/RS Study Guide: A Guide for Environmental Health Responsibilities and Competencies (5th edition) National Environmental Health Association (NEHA),2024-01-01 The REHS RS Study Guide reflects the most recent changes and advancements in environmental health technologies and theories Incorporating the insights of 29 subject matter experts from across academia industry and the regulatory community paired with references from over 30 scholarly resources this essential reference is intended to help those seeking to obtain the NEHA Registered Environmental Health Specialist Registered Sanitarian Credential Table of Contents General Environmental Health Statutes and Regulations Food Protection Potable Water Wastewater Solid and Hazardous Waste Hazardous Materials Zoonoses

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Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993

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