

Student Exploration: pH Analysis

Essential Question: How do acids and bases differ in their properties?

Background: Acids and bases are two types of chemical compounds that can be found in many everyday substances.

- Acids** are substances that produce hydrogen ions (H^+) when dissolved in water. Some common examples of acids are:
 - Hydrochloric acid (stomach acid)
 - Vinegar (acetic acid)
 - Citric acid (found in lemons)
 - Sulfuric acid (found in car batteries)
- Bases** are substances that produce hydroxide ions (OH^-) when dissolved in water. Some common examples of bases are:
 - Sodium hydroxide (lye)
 - Ammonia (found in cleaning products)
 - Calcium hydroxide (found in cement)
 - Sodium carbonate (found in baking soda)

Activity Overview

The purpose of this activity is to determine the pH of various substances. The pH scale is a logarithmic scale that ranges from 0 to 14. Substances with a pH of 7 are neutral, substances with a pH below 7 are acidic, and substances with a pH above 7 are basic. The pH of a substance can be determined using a pH meter or pH paper.

The pH of a substance is a measure of its acidity or basicity. The pH scale is a logarithmic scale that ranges from 0 to 14. Substances with a pH of 7 are neutral, substances with a pH below 7 are acidic, and substances with a pH above 7 are basic.



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