ACIDS, BASES AND SALTS

Definition of an acid	What causes the acidic properties of acids?		
An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	The hydrogen ions (H') cause the acidic properties and these are formed in the presence of water.		

Another term to refer to an acid	Why an acid is also called a proton donor?	
An acid is called a <u>proton donor</u>	It's because an acid provides protons or hydrogen ions (H·) to other	
	substances during the reaction.	

Substance to which an	A base is then called	Why the base is called a	Equation for the reaction between
acid provides protons		proton acceptor?	acid and base
BASE	PROTON ACCEPTOR	Because the base accepts hydrogen ions from acids	H [*] _{jeg} * ÖH _{jeg} * H _j O _R

Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?	
Hydrochloric acid	HC/		They are derived from mineral salts ie chlorides for	
Sulphuric acid	H;80.	MINERAL ACIDS	HCl, sulphates for H ₂ SO ₄ and nitrates for HNO ₂	
Nitric acid	HNO ₃	Fig. 200 Resident States and C		

Other minera known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H ₂ SO ₃	Derived from sulphites	Ethanoic acid	CITRIC ACID from lemons
Carbonic acid	H ₂ CO ₂	Derived from carbonates	(CH ₃ COOH)	TARTARIC ACID from grapes
Phosphoric acid	H ₂ PO ₄	Derived from phosphates		ACETIC ACID from vinegar
Nitrous acid	HNO ₃	Derived from nitrites	Methanoic acid	LACTIC ACID from sour milk
DUMPARIES :	Production of the	Property of the state of the st	(HCOOH)	Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces	Term given to the number of hydrogen ions produced by one molecule of an acid	Definition of basicity of an acid
HYDROGEN IONS	BASICITY OF AN ACID	BASICITY of an acid is the <u>number of</u> <u>hydrogen ions</u> produced by <u>one</u> <u>molecule</u> of an acid <u>in aqueous</u> <u>solution</u> .

Study Guide Acids And Bases

James Adams, Maryn Munley, Nathan
Eisenbeis, Caroline Miller, Jesse Shetler

Study Guide Acids And Bases:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms. The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in The Practice of Chemistry Study Guide & Solutions Manual Pamela living systems and the physical world around us Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises **Study Guide for Fundamentals of Nursing** E-Book Barbara L. Yoost, Lynne R. Crawford, Patricia Castaldi, 2019-02-28 Coordinated chapter by chapter to the new second edition of Yoost and Crawford's Fundamentals of Nursing this engaging study tool encompasses a wealth of activities and features like case study questions study group challenges self review questions and more to keep your learning interactive meaningful and fun For chapters that include skills detailed and interactive skills checklists will be provided on the Fundamentals of Nursing Evolve website Answers to all questions will be included in the back of the study guide to help you assess your learning Chapter review questions include multiple choice matching true false completion image labelling and

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