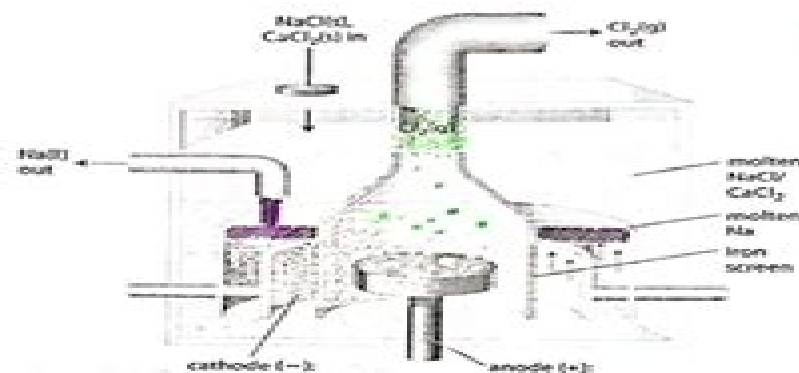


IGCSE – HONORS – ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al_2O_3)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al^{3+} , O^{2-}

Half equations:

In the Cathode $(Al^{3+} + 3e^- \rightarrow Al) \times 4$

In the Anode $(2O^{2-} \rightarrow O_2 + 4e^-) \times 3$

Overall Equation $4Al^{3+} + 6O^{2-} \rightarrow 4Al + 3O_2$

Observations:

Cathode: The Molten Aluminum can be observed

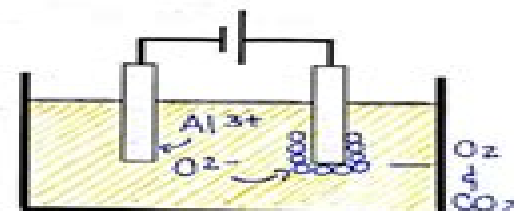
Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O_2 to form CO_2

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



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